

MagOzone

Magnesium information

4) The complex oxides of magnesium in *MagOzone*, are not for nutritional use but act as a cleanser drawing water into the digestive tract which supports the cleansing effect.

“A common use for high doses of oral magnesium salts is to produce a laxative effect to treat constipation. In the intestinal lumen the poorly absorbable magnesium ions (and other ions such as sulphate) exert an osmotic effect and cause water to be retained in the intestinal lumen. This increases the fluidity of the intraluminal contents and results in a laxative action. Although the laxative action of magnesium is thought to be due to a local effect in the intestinal tract, it is also possible that released hormones such as cholecystokinin or activation of constitutive nitric oxide synthase might contribute to this pharmacological effect. Under normal circumstances the pharmacological administration of high doses of oral magnesium salts is safe and some salts--such as magnesium hydroxide--also have an antacid effect to neutralize stomach acid. However, high doses of magnesium or prolonged use may allow sufficient absorption into the systemic circulation to cause renal or other organ toxicity.” <https://www.ncbi.nlm.nih.gov/pubmed/8878010>

Oxides of Magnesium have been formulated to hold and release oxygen over an extended amount of time.

“The most common oxygen-releasing materials include sodium percarbonate (4), calcium peroxide (3, 9), magnesium peroxide (4), hydrogen peroxide (10, 11), and fluorinated compounds (12-14). Solid peroxides decompose upon exposure to water to release oxygen. However, if this process takes place too quickly, it may significantly damage the cells due to free radical formation (15). The rate of oxygen release via peroxide compounds is influenced by a number of factors such as, temperature, pH, and presence of a buffer or catalyst (16, 17). For example, when solid peroxide compounds react with water they form metal hydroxides, which induce an increase in the pH and the amount of released oxygen (18). Alternatively, the use of buffers provide adjustment of pH and therefore oxygen generation. Moreover, the purity and solubility of peroxides significantly affect the kinetics of oxygen release.”

<https://www.ncbi.nlm.nih.gov/pmc/articles/PMC3708668/>